Dataset on constructing colloidal nanoparticles into dry nano-micro-particle (NMP) powders with nanoscale magnetic, plasmonic and catalytic functionalities


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Data Article

Dataset on constructing colloidal nanoparticles into dry nano-micro-particle (NMP) powders with Nanoscale Magnetic, Plasmonic and Catalytic Functionalities

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Abstract

The data presented in this article is related to the research article entitled “A One-Pot Method for Building Colloidal Nanoparticles into Bulk Dry Powders with Nanoscale Magnetic, Plasmonic and Catalytic Functionalities” (Ye et al., 2019). The data shows the hydrophobicity of the nanoparticle (NP) building blocks used for constructing NMPs obtained through contact angle measurements, along with the effect of NP hydrophobicity on the stability of the parent Pickering emulsions. SEM data of the morphology of NMPs is presented. Finally, a mathematical model is presented to predict the average diameter of NMPs produced via different experimental parameters.

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1. Data

Fig. 1 compares the hydrophobicity of a surface composed of different types of densely packed NPs as well as the same type of NP but with different surface-cappings. Table 1 presents a list of capping agents along with the corresponding amounts required for modifying the surface of Au NPs to achieve the correct hydrophobicity to generate stable Pickering emulsions. Fig. 2 show SEM images of NMPs with different morphologies obtained by tuning the polystyrene concentration. Fig. 3 presents SEM data of the surface morphology of NMPs before and after removal of the NP layer on the surface. A mathematical model and supporting mathematical data is provided in Fig. 4 to illustrate the effect of experimental parameters on the size of product NMPs.

2. Experimental design, materials and methods

2.1. Surface hydrophobicity data for NMPs

Contact angle measurements were performed on surface exposed nanoparticle sheets (SENS) fabricated following a method previously reported by our group [2,3]. SENSs are formed by fixing a monolayer 2D NP array assembled using promoters at the liquid-liquid interface (LLI) onto a thin polymer sheet in-situ through solvent evaporation. Therefore, SENSs have the same micro-structure as NMPs which makes them perfect candidates for contact angle measurements [1]. As shown in the contact angle data, citrate-capped NPs (θ = 75°) were notably more hydrophobic than PVP-capped NPs (θ = 38°) which generated stable Pickering emulsions. A full list of hydrophilic capping agents capable of generating stable Pickering emulsions with the aid of promoters tested in this work is shown in Table 1.
Calculations of the amount of molecular capping needed for Au particle surface modification:

The size of Au particles as determined by SEM images varied from 28 nm to 34 nm, therefore the average size of the Au particles is 30 nm in diameter. The surface occupied by one modifier molecule is 0.20 nm² [4].

The amount of Au used is:

\[
\frac{\text{Molecular mass of Au} \times \text{Mass of Au chloride}}{\text{Molecular mass of Au chloride}} = \frac{197 \times 0.005}{394} = 0.0025 \text{ g}
\]

The volume of one Au NP is:

\[
\frac{4}{3} \pi R^3 = \frac{4}{3} \pi (15)^3 = 1.41 \times 10^4 \text{ nm}^3
\]

The mass of one Au NP is:

\[
\rho V = \left(19.30 \times 10^{-21}\right) \times \left(1.41 \times 10^4\right) = 2.73 \times 10^{-16} \text{ g}
\]

Fig. 1. Contact angle data for SENS carrying different types of NP layers or the same type of NP with different types of molecular capping. SEM data of the surface of typical SENS films. All scale bars in SEM images correspond to 100 nm. All values shown are based on automated droplet profile fitting and are accurate within ± 5°.
The number of Au NP in 5 mL colloid is:

\[
\frac{\text{Amount of Au used}}{\text{Mass of one Au NP}} = \frac{0.0025}{2.73 \times 10^{-16}} = 9.16 \times 10^{12}
\]

The surface area of one Au NP is:

\[
= 4\pi R^2 = 4\pi (15 \text{ nm})^2 = 2827.43 \text{ nm}^2
\]

The total surface area in 5 mL colloid is:

\[
= \text{Number of Au NP } \times \text{ surface area of one NP }
\]
\[
= 9.16 \times 10^{12} \times 2827.43 = 2.5916 \times 10^{16} \text{ nm}^2
\]

The amount of modifier required for surface functionalization is:

\[
= \frac{\text{Total surface area in 5 mL Au colloid}}{\text{Surface area of one modifier molecule}}
\]

<table>
<thead>
<tr>
<th>Chemical name</th>
<th>Chemical structure</th>
</tr>
</thead>
<tbody>
<tr>
<td>3-Mercaptotetronic acid (MPA)</td>
<td>![MPA structure]</td>
</tr>
<tr>
<td>Thioglycolic acid (TGA)</td>
<td>![TGA structure]</td>
</tr>
<tr>
<td>2-Mercapto-5-benzimidazolesulfonic acid sodium salt (MBS)</td>
<td>![MBS structure]</td>
</tr>
<tr>
<td>3-Mercapto-1-propanesulfonate (MPS)</td>
<td>![MPS structure]</td>
</tr>
<tr>
<td>Polyvinylpyrrolidone (PVP)</td>
<td>![PVP structure]</td>
</tr>
</tbody>
</table>
The amount of modifier used is:

\[
\text{amount} = \frac{2.59 \times 10^{16}}{0.20} = 1.30 \times 10^{-7} \text{ mol}
\]

\[
\text{Concentration of modifier} \times \text{volume of modifier} = 10^{-2} \times 10^{-4} = 1 \times 10^{-6} \text{ mol}
\]

\(\rho, V, R\) refer to density, volume and radius, respectively.

---

**Fig. 2.** SEM data of NMPs observed at different stage of buckling caused by DCM evaporation: (a) smooth; (b) buckyball; (c) deformed; (d) labyrinth. All scale bars in (a–d) correspond to 10 \(\mu\)m. Insets in (a) and (d) show typical high magnification SEM images of the NP surface layers on the NMPs. All scale bars in inset correspond to 100 nm. (e–g) show images of NMPs made from 0.2 g/mL, 0.08 g/mL and 0.02 g/mL polystyrene/DCM concentrations, respectively. All scale bars in (e–g) correspond to 100 \(\mu\)m.

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**Fig. 3.** SEM data showing the surface of a magnetic TiO\(_2\) NMP before (a) and after sonication (b). All scale bars correspond to 100 nm.
2.2. SEM data of NMPs

As shown in the SEM data in Fig. 2a–d, the morphology of the NMPs becomes more and more irregular as the concentration of the polystyrene decreases. Fig. 2e and f shows the relative population of deformed NMPs increasing with decreasing concentrations of polystyrene.

As shown in Fig. 3a, prior to sonication, a densely packed layer of TiO2 NPs are present on the surface of the NMP. After extensive sonication, the NPs on the surface are removed to show the bare and smooth surface of the micro-polymer core.

2.3. NMP size distribution data

Fig. 4a–c compares the size of the parent Pickering emulsions with their corresponding NMP products. A surface plot illustrating the effect of polymer concentration and colloid-to-oil ratio is shown in Fig. 4d. Within the plot, the lower surface (yellow) represents ideal smooth and spherical NMPs and the offset surfaces (red and blue) represent the effect of different extents of wrinkling to the final diameter of the NMPs. In general, the data shows that when the effects of wrinkling and buckling are taken into account, the predicted size of NMPs becomes larger. Detail mathematical data correlating to the size of product NMPs versus colloid-oil ratio and polystyrene concentration is shown below:

![Graph of NMP size distribution data](image-url)
Assuming the number of Au particle is $n$, the volume of the oil is $v\ \mu m^3$, the number of the emulsion droplets is $a$, the concentration of the polymer in the oil is $c \ g/mL \ (b \ g/\mu m^3)$, the radius of one emulsion droplet is $r \ \mu m$, the radius of the NMPs is $R \ \mu m$. Assuming the Au NPs are spherical, and the diameter of the Au particles $d$ is 0.03 $\mu m$. The density of polystyrene $\rho$ is 1.04 $g/mL \ (1.04 \times 10^{-12} \ g/\mu m^3)$.

The surface area of an Au NP is:

\[
= \frac{\pi d^2}{4} \ \mu m^2
\]

The whole surface area of all Au NPs is:

\[
= \frac{\pi d^2}{4} n \ \mu m^2
\]

The volume of one emulsion droplet is:

\[
= \frac{v}{a} \ \mu m^3
\]

The radius of one emulsion droplet is:

\[
= \frac{3v}{4\pi a} \ \mu m
\]

The surface area of one emulsion droplet is:

\[
= 4\pi r^2 = 4\pi \left(\frac{3v}{4\pi a}\right) \ \mu m^2
\]

The surface area of the all emulsion droplets equals to the surface area of all Au NPs:

\[
4\pi r^2 a = 4\pi a \left(\frac{3v}{4\pi a}\right) = \frac{\pi d^2}{4} n
\]

From these equation one derives the volume of one emulsion droplet:

\[
\frac{v}{a} = \frac{d^3}{233\nu^{0.5}} (\frac{v}{n})^{1.5} = \frac{233}{d^6} \left(\frac{v}{n}\right)^3 \ \mu m^3
\]

When $v = 10^{12} \ \mu m^3$, $n = 10 \times 10^{12}$, the radius of the emulsion droplet $R$ is calculated to be 0.42 mm, which is close to the experimental value 0.25—0.4 mm.

The mass of polystyrene in one emulsion droplet is:

\[
\frac{vb}{a} = \frac{233}{d^6} \left(\frac{v}{n}\right)^3 b \ g
\]

The volume of polystyrene in one emulsion droplet is:
\[ = 2.24 \cdot 10^{14} \frac{d^6}{v_n^3} \frac{b \mu m^3}{\mu m^3} \]

The radius of the NMP is:

\[ = \frac{37677}{d^2} \frac{v}{n} \sqrt{b} \mu m \]

\[ = \frac{3.77}{d^2} \frac{v}{n} \sqrt{c} \mu m \]

When \( v = 10^{12} \mu m^3, n = 10 \times 10^{12}, \) the radius of the emulsion droplet \( R \) is calculated to be 180 \( \mu m \), which is quite different from the observed value 50 \( \mu m \).

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**Transparency document**

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**References**


